



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

CENTRE
NUMBER

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CANDIDATE
NUMBER

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COMBINED SCIENCE

0653/03

Paper 3 (Extended)

October/November 2008

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
1	
2	
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7	
8	
9	
Total	

This document consists of **18** printed pages and **2** blank pages.



1 Fig. 1.1 shows a food web.

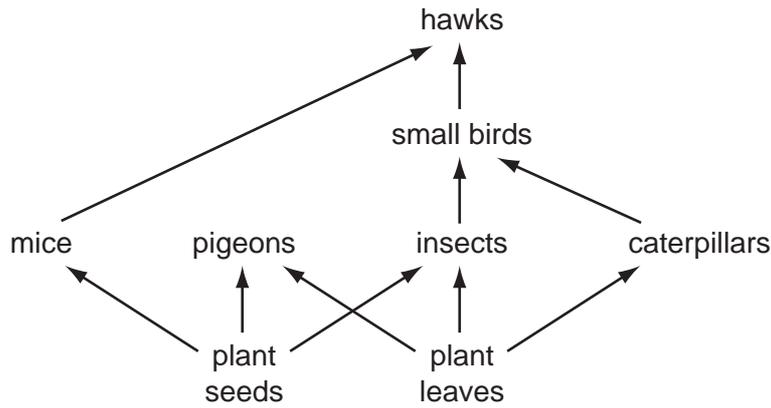


Fig. 1.1

(a) (i) State what the arrows in Fig. 1.1 represent.

..... [1]

(ii) The longest food chain in Fig. 1.1 has four organisms. Explain why it is rare for food chains to be longer than this.

.....
.....
..... [2]

(b) Describe how an atom of carbon in a glucose molecule in an insect could become part of a glucose molecule in a plant leaf.

.....
.....
.....
.....
..... [4]

- 2 (a) Two inflated rubber rings, one black and one white, are left on a hot beach in the

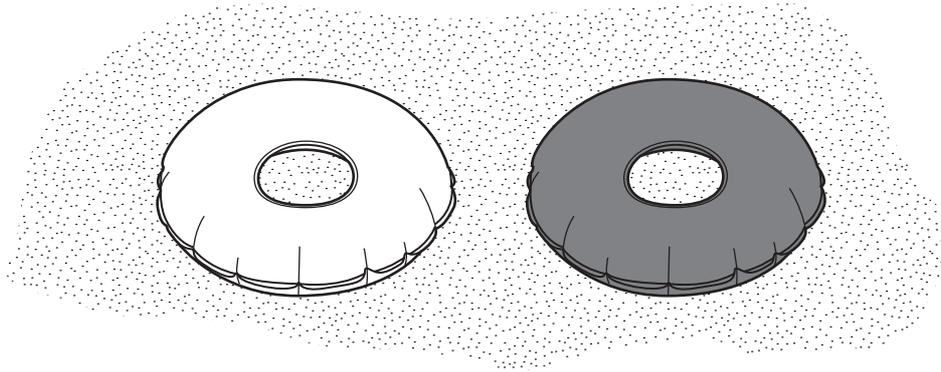


Fig. 2.1

Explain why the temperature of the air inside the black rubber ring rises more quickly than that in the white rubber ring.

.....

.....

..... [2]

- (b) Someone has left a glass bottle on the beach. The curved glass acts like a lens focussing the sun's rays.

Complete the light rays on Fig. 2.2 to show what happens to rays of light after they have passed through a convex lens.

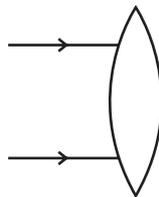


Fig. 2.2

[2]

(c) The power of the waves is used as a renewable source of energy.

Suggest how the motion of the waves can be converted to electrical energy.

.....
.....
..... [2]

(d) A girl collects two pebbles of the same size from the beach but one seems to be heavier than the other. How could she measure the densities of the two pebbles?

.....
.....
.....
..... [3]

- 3 The chemical symbol of magnesium is shown below.



- (a) Draw a labelled diagram of an atom of magnesium.

Your diagram should show the numbers of nucleons and the electron configuration.

[2]

- (b) Magnesium is produced industrially by the electrolysis of molten magnesium chloride. Fig. 3.1 shows a simplified diagram of this process.

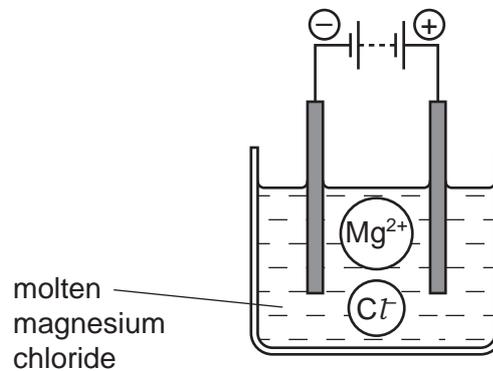


Fig. 3.1

- (i) Describe, in terms of ions and electrons, what happens at the surface of the cathode.

.....

.....

..... [2]

(ii) Use the information in Fig. 3.1 to explain why the chemical formula of magnesium chloride is $MgCl_2$.

.....
..... [1]

(c) A student added magnesium to dilute hydrochloric acid as shown in Fig. 3.2.

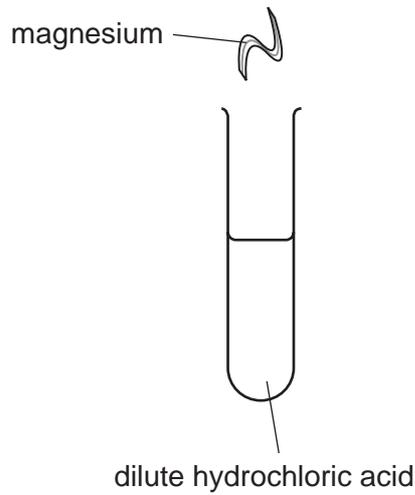


Fig. 3.2

(i) Explain, in terms of electrons, why the magnesium atoms in the reaction in Fig. 3.2 are said to be oxidised.

.....
.....
..... [2]

(ii) Explain, in terms of ions, why the pH of the mixture in Fig. 3.2 increases when magnesium is added to the acid.

.....
.....
..... [2]

4 Fig. 4.1 shows part of the male reproductive system.

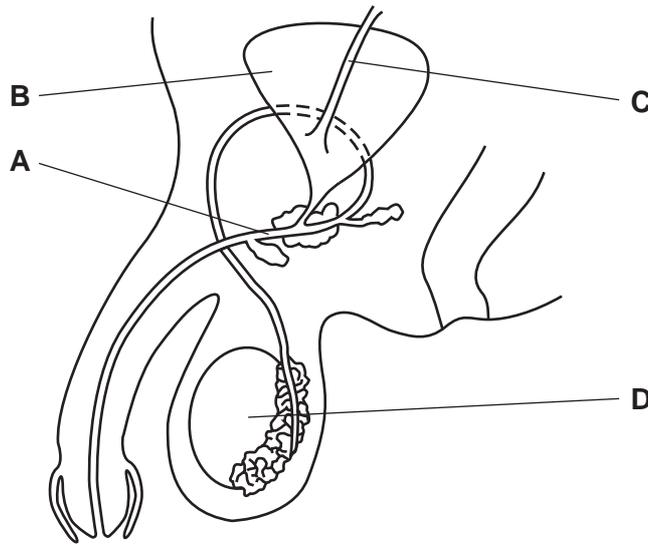


Fig. 4.1

(a) (i) Name the parts labelled **A** and **B**.

A

B

[2]

(ii) Describe the functions of parts **C** and **D**.

C

.....

D

..... [2]

(iii) On Fig. 4.1, write the letter **X** to show the part of the reproductive system which is cut or tied when a man has a sterilisation operation. [1]

(b) Humans reproduce using sexual reproduction. Sexual reproduction produces offspring that are genetically different from each other and from their parents.

Explain how this can be an advantage to a species of organism.

.....

.....

.....

..... [3]

- 5 (a) Fig. 5.1 shows two cars **A** and **B**.

Car **A** produces exhaust gases which appear black. The exhaust gases from car **B** cannot be seen. Both cars have engines which use diesel (gas oil) which is a hydrocarbon fuel.

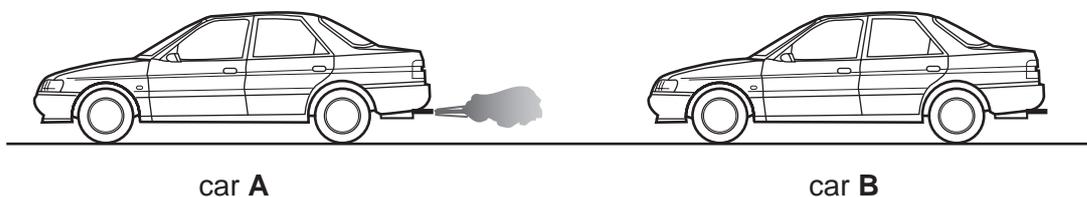


Fig. 5.1

- (i) Describe briefly how hydrocarbon fuels like diesel are obtained.

.....

 [2]

- (ii) The formula of a typical molecule in diesel is $C_{13}H_{28}$.

Calculate the relative molecular mass, M_r , of $C_{13}H_{28}$.

Show your working.

..... [2]

- (b) The energy needed to move cars is provided by the combustion of the fuel. Air is supplied to the engine for this combustion to occur.

Fig. 5.2 shows a bar chart of the main gases in a sample of dry air.

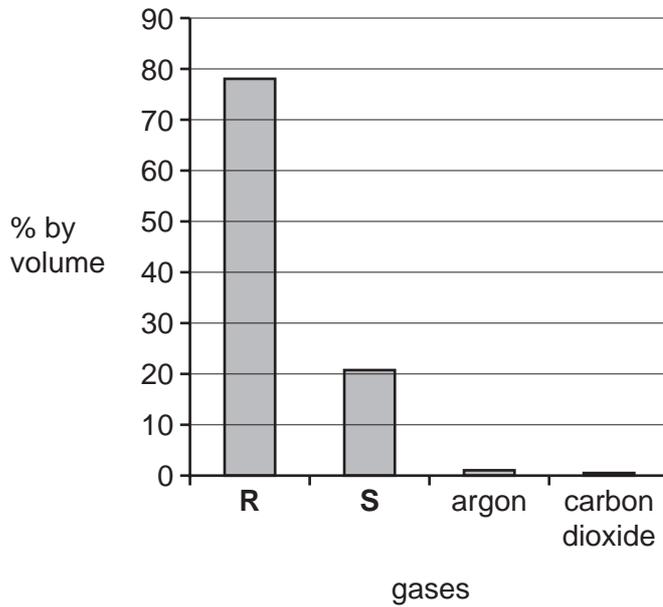


Fig. 5.2

- (i) Explain which gas shown in Fig. 5.2 reacts with the diesel fuel in car engines.

.....
 [1]

- (ii) Suggest the name of the black substance in the exhaust gases of car **A** in Fig. 5.1, and explain briefly how it is formed.

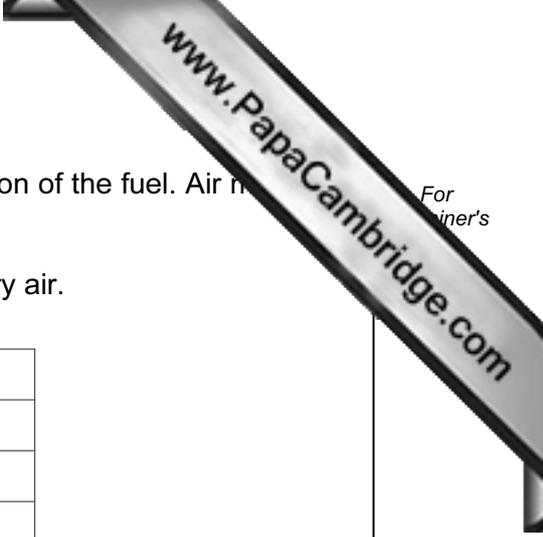
.....

 [2]

- (iii) Explain why car engines should never be left running for long periods of time in a garage or other enclosed space where there are people.

.....

 [2]



- 6 (a) The isotope radon-220 is radioactive. A sample was investigated to find its half-life. The activity of the isotope was measured every 30 seconds for 6 minutes. The results are shown in Fig. 6.1.

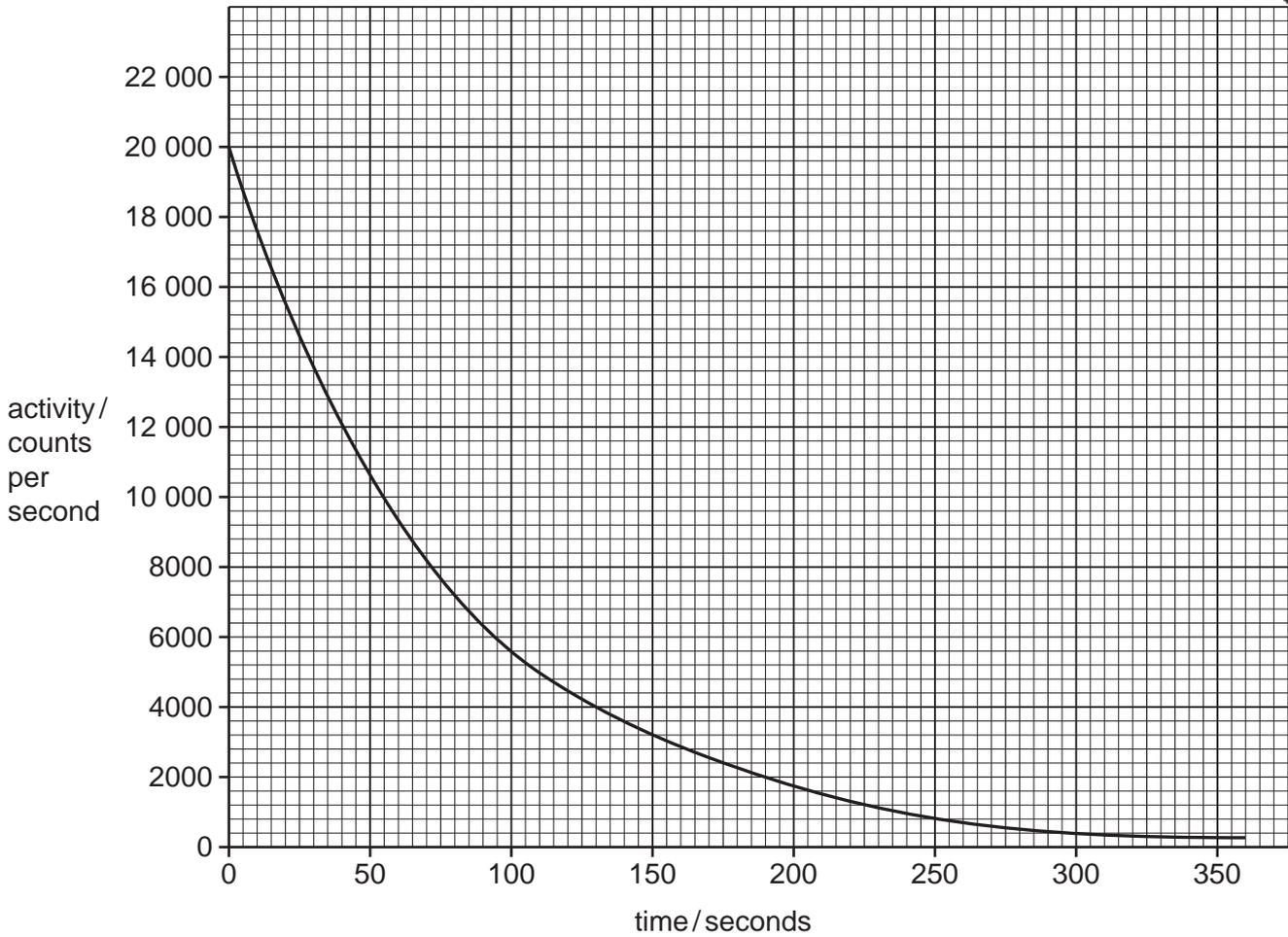
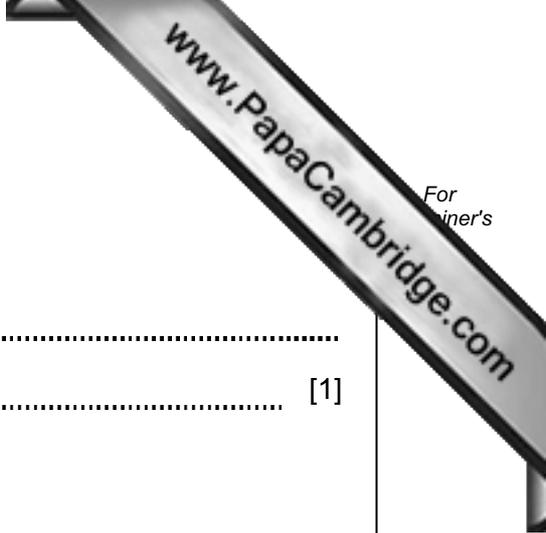


Fig. 6.1

Use the graph to calculate the half-life of the isotope. Show your working on the graph.

..... [2]



(b) There are several isotopes of radon.

State the meaning of the word *isotope*.

.....
..... [1]

(c) Radon-220 emits alpha radiation.

Explain why alpha radiation is dangerous to human beings.

.....
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.....
.....
.....
.....
..... [4]

7 This article appeared in a newspaper in Pakistan in 2006.

Many more people in Pakistan and India are developing diabetes. This is an illness where the regulation of blood glucose does not work properly. It is dangerous because rising levels of glucose in the blood can damage cells in many parts of the body, including the blood system and the eyes.

Doctors think that the increase in diabetes is happening because people are eating more fast food. Where they used to eat a lot of rice and lentils, they are now eating more fried foods and greasy take-aways. As well as increasing the risk of diabetes, it is causing an increase in obesity. This also increases the risk of heart disease.

(a) The regulation of blood glucose is part of homeostasis.

Explain the meaning of the term *homeostasis*.

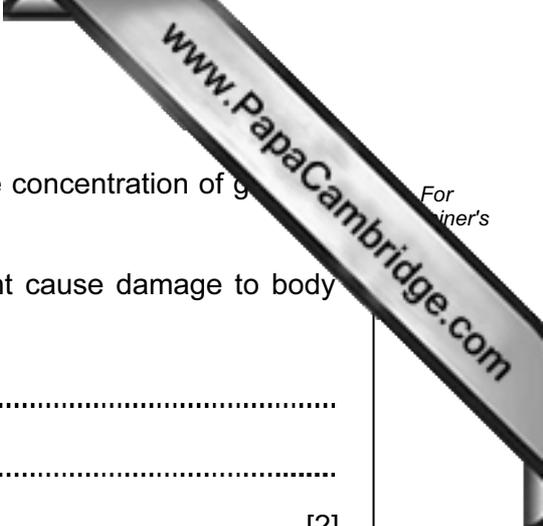
.....
..... [2]

(b) (i) Name the hormone that is produced when the blood glucose level rises, and which helps to bring it back down to normal.

..... [1]

(ii) Describe how the hormone reduces the amount of glucose in the blood.

.....
.....
..... [2]



(c) When a person with diabetes eats food containing sugar, the concentration of glucose in their blood increases.

Using what you know about osmosis, explain how this might cause damage to body cells.

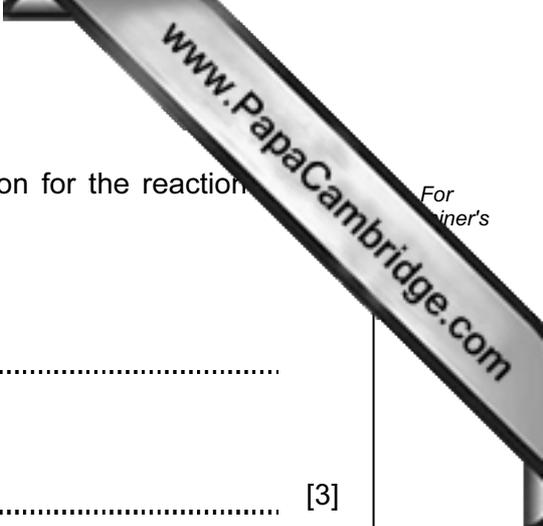
.....
.....
..... [2]

(d) (i) Suggest why eating foods containing a lot of fat, rather than eating lentils and rice, can lead to obesity.

.....
.....
..... [2]

(ii) Explain how a poor diet can increase the risk of a heart attack.

.....
.....
.....
..... [2]



- (i) Write a word equation and a balanced symbolic equation for the reaction that occurs when calcium carbonate is heated strongly.

word equation

.....

symbolic equation

..... [3]

- (ii) Name solution **S** in Fig. 8.2, and predict what would be observed during the reaction.

.....
..... [1]

- (iii) Describe how the student could test the solid which remains in tube **T** to find out if all the calcium carbonate had reacted.

.....
.....
.....
..... [2]

- (iv) Why are large amounts of calcium carbonate sometimes spread on soil which is going to be used for growing crops?

.....
..... [1]

9 (a) A student has six resistors as shown in Fig. 9.1.

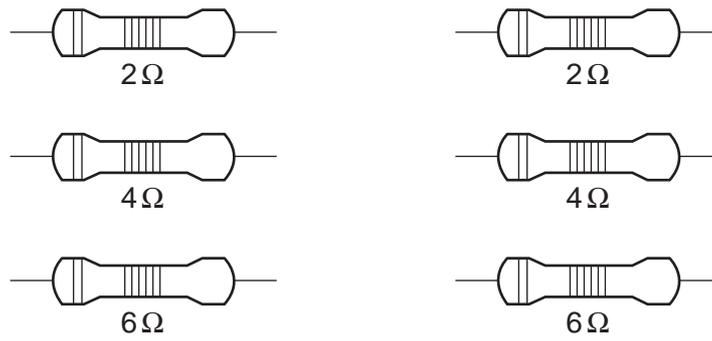


Fig. 9.1

(i) Describe how he can combine **two** of these resistors to get a total resistance of 10 ohms.

.....
 [1]

(ii) Explain how he can combine **two** of these resistors to get a total resistance of three ohms.

.....
 [3]

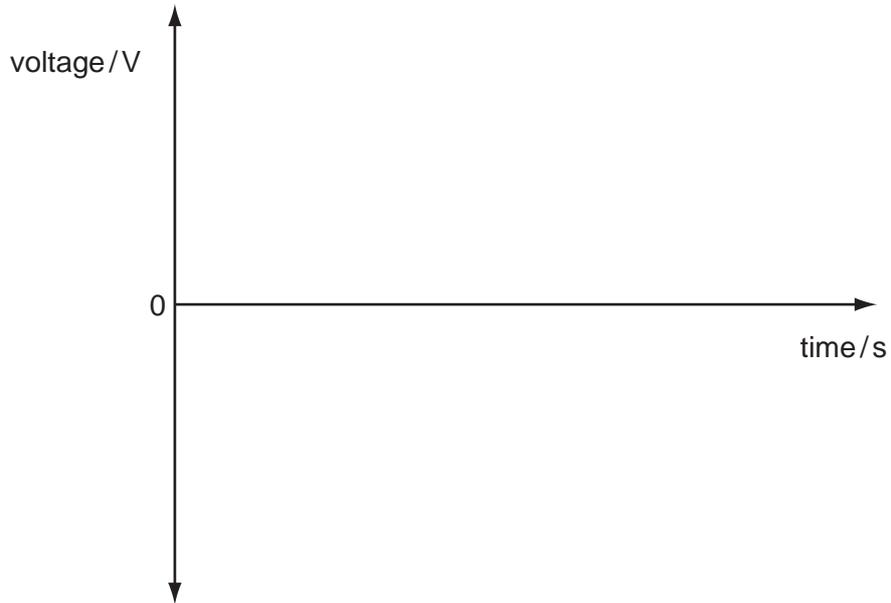
(b) Electricity can be generated by turning a coil of copper wire in a magnetic field.

(i) Describe **two** ways to increase the voltage produced by this generator.

1
 2 [2]

- (ii) Generators can supply an alternating current which has a frequency of 50 Hz.

On the grid below, sketch a graph to show the current produced by this alternating current generator during a period of 0.1 seconds.



[3]

- (c) Electricity is often transmitted through overhead power cables hung from pylons. If these cables are put up on a hot summer day, they are hung loosely from the pylons as shown in Fig. 9.2.

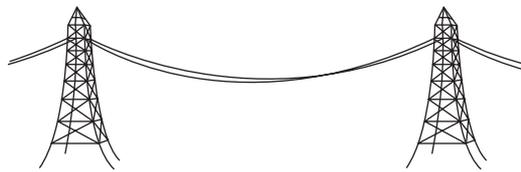


Fig. 9.2

Suggest why they are hung loosely.

.....
.....
.....

[2]

DATA SHEET
The Periodic Table of the Elements

		Group															
I	II	III	IV	V	VI	VII	0							0			
		1 H Hydrogen 1											2 He Helium 2				
7 Li Lithium 3	9 Be Beryllium 4											20 Ne Neon 10					
23 Na Sodium 11	24 Mg Magnesium 12	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9							35.5 Cl Chlorine 17				
39 K Potassium 19	40 Ca Calcium 20	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17							40 Ar Argon 18				
85 Rb Rubidium 37	88 Sr Strontium 38	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35							84 Kr Krypton 36				
133 Cs Caesium 55	137 Ba Barium 56	65 Zn Zinc 30	64 Cu Copper 29	59 Ni Nickel 28	56 Fe Iron 26	59 Co Cobalt 27	64 Cu Copper 29	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	127 I Iodine 53			131 Xe Xenon 54
226 Ra Radium 88	227 Ac Actinium 89	204 Tl Thallium 81	201 Hg Mercury 80	195 Pt Platinum 78	190 Os Osmium 76	192 Ir Iridium 77	197 Au Gold 79	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84			210 Rn Radon 86
												169 Tm Thulium 69	175 Lu Lutetium 71				
												167 Er Erbium 68	173 Yb Ytterbium 70				
												166 Fm Fermium 100	168 No Nobelium 102				
												165 Ho Holmium 67	171 Yb Ytterbium 70				
												162 Dy Dysprosium 66	173 Yb Ytterbium 70				
												159 Tb Terbium 65	173 Yb Ytterbium 70				
												157 Gd Gadolinium 64	173 Yb Ytterbium 70				
												152 Eu Europium 63	173 Yb Ytterbium 70				
												150 Sm Samarium 62	173 Yb Ytterbium 70				
												144 Nd Neodymium 60	173 Yb Ytterbium 70				
												141 Pr Praseodymium 59	173 Yb Ytterbium 70				
												140 Ce Cerium 58	173 Yb Ytterbium 70				
												238 U Uranium 92	173 Yb Ytterbium 70				
												232 Th Thorium 90	173 Yb Ytterbium 70				
												94 Pu Plutonium 94	173 Yb Ytterbium 70				
												93 Np Neptunium 93	173 Yb Ytterbium 70				
												91 Pa Protactinium 91	173 Yb Ytterbium 70				
												96 Cm Curium 96	173 Yb Ytterbium 70				
												95 Am Americium 95	173 Yb Ytterbium 70				
												97 Bk Berkelium 97	173 Yb Ytterbium 70				
												98 Cf Californium 98	173 Yb Ytterbium 70				
												99 Es Einsteinium 99	173 Yb Ytterbium 70				

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X
= relative atomic mass	
= atomic symbol	
= proton (atomic) number	

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).